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| Unit 2- Ch 9 Naming Chemicals Chem111/112 |

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| \*\*How do you name? Make up a page with an example for each category.\*\* |

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| 1.Binary Ionic?  2.Ternary Regular?  3.Ternary Irregular? | Hydrates? | Bases |  | 1.Binary Acids  2.Ternary Regular Acids  3.Ternary Irregular? | Molecular  1. Diatomics  2. Binary  3. Polyatomics  4. List ones that you had to study  ex. H2O2 |  |

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| Give some diagnostic test used to distinguish between acids and Bases. |

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| Fill in the missing state or formula or name. |

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| Fomula(state) | Name | Formula(state) | Name: |
| 1. | Sodium sulfate deca hydrate | 18 Fe2O3(s) |  |
| 2 NH4NO3(aq) |  | 19 | nitrogen |
| 3 | phosphorous trihydride | 20 SO3(g) |  |
| 4 H3PO5(aq) |  | 21 | ammonium chloride |
| 5 | silicon dioxide | 22 P4(s) |  |
| 6 HF(aq) |  | 23 | carbon tetrachloride |
| 7 N2H4(g) |  | 24 C2H5OH(l) |  |
| 8 | perchloric acid | 25 | Methane |
| 9 | nickel(II) bromide | 26 | acetic acid |
| 10 PbO2(s) |  | 27 | nitrous acid |
| 11 | tetra phosphorous decaoxide | 28 HCl(aq) |  |
| 12 KNO3(aq) |  | 29 HClO4(aq) |  |
| 13 | sucrose | 30 KCl(aq) |  |
| 14 | propane | 31 | calcium carbonate |
| 15 SiO2(s) |  | 32 C(s) |  |
| 16 CaSO2●2H2O(s) |  | 33 | Iodine |
| 17 | Lead(IV)oxide | 34 FeSO4(s) |  |

🡪Using a different colour pen write i-ionic, a-acid, b-base, m-molecular in front of each formula above. More on next page.

🡪Attach a page folded in four with the properties, diagnostic tests, states and how to recognize each category of chemical from its formula. Write all of the formulas from above into your table.

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| Ionic Ex.FeSO4(s)  Properties list others  Diagnostic tests from above  State  How to recognize? |  |
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